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titrations key, Work 23 strong acidstrong base titrations, Work 26, Titrations work, Titrations practice work. Click on pop-out icon or print icon to worksheet to print or download.

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In a strong acid - strong base titration, neutralization produces water and an aqueous solution of a salt, whose cation and anion come from the base and acid, respectively. Neither ion is acidic or basic, so the pH at the equivalence point is that of neutral water; i.e., 7.00.

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Titrations worksheet W 336 Everett Community College Tutoring Center Student Support Services Program 1) It takes 83 mL of a 0.45 M NaOH solution to neutralize 235 mL of an HCl solution. What is the concentration of the HCl solution? 2) You are titrating an acid into a base to determine the concentration of the base.

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Worksheet 22 - Weak Acid/Strong Base Titrations A. Initial pH
This is determined by the initial concentration of the weak acid and the dissociation constant, K_a .
 $[H^+][A^-] / [HA] = K_a$
Set up an ICE table for the dissociation of the weak acid in water.
The "change" line

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Worksheet 23 - Strong Acid/Strong Base Titrations A. Initial pH
This is always determined based solely on the initial concentration of the acid or base being titrated. Every mole of acid or base will produce one mole of H_3O^+ or OH^- . Exception: Bases formed with Group II cations will form two moles of OH^- for every mole of base, e.g. $Ca(OH)_2$.

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