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This activity demonstrates the links between the topics of rates of reaction and the equilibrium law. It provides students with an explanation of the equilibrium law and helps them explain why Le Chatelier's principle works for temperature, concentration and

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Objectives. After completing this section, you should be able to write the equilibrium constant expression for a given reaction. assess, qualitatively, how far a reaction will proceed in a given direction, given the value of K_{eq} ; explain the difference between rate and equilibrium.

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$r = -\frac{1}{a} \frac{dC_A}{dt} = -\frac{1}{b} \frac{dC_B}{dt} = \dots$ from the expression above, it is clear that the usual convention is that reaction rate is taken as products formation rate. Rate is a function of the concentration of reactants and products, temperature, pressure and presence of a catalyst.

6.7: Describing a Reaction - Equilibria, Rates, and Energy ...

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Rates and equilibria of organic reactions as treated by statistical, thermodynamic, and extrathermodynamic methods. New York : Dover Publications, 1989, ©1963 (DLC) 89011941

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Details The quantitative prediction of rates and equilibria of reactions is one of the most important and most difficult tasks of chemistry.

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II QUANTITATIVE CORRELATION OF RATES AND EQUILIBRIA

References: 1. Carey and Sundberg, Part A, pp. 196-209 (3rd ed.) (4th Ed. pp. 204-215) (5th Ed. pp. 335-344) 2. Carroll, p. 371 3. Miller, p.122 4. L&R pp. 143-158 5. Isaacs Chapter 4 6. Anslyn/Dougherty, pp 445-451 6. J. Hine, Structural Effects on Equilibria in Organic Chemistry (see p.1) 7.

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